

Name: _____

Ionic Compounds—Univalent Metal Ions

1. If the following pairs of elements were mixed and heated, they would combine into solid ionic compounds. Write the name and formula of each compound formed.

	Name	Formula
a) silver and iodine	silver iodide	$\text{AgI}_{(s)}$
b) magnesium and oxygen	_____	_____
c) magnesium and bromine	_____	_____
d) calcium and nitrogen	_____	_____
e) zinc and selenium	_____	_____
f) sodium and sulfur	_____	_____
g) barium and phosphorus	_____	_____
h) aluminium and fluorine	_____	_____
i) potassium and chlorine	_____	_____
j) silver and oxygen	_____	_____

continued...

Unit A—Energy and Matter in Chemical Change

2. Write the correct names for each of the following compounds.

a) MgCl_2 _____

b) Ag_3N _____

c) CsF _____

d) CdO _____

e) MgBr_2 _____

f) Al_2O_3 _____

g) NaI _____

h) K_2S _____

i) BaS _____

j) Li_3P _____

Unit A—Energy and Matter in Chemical Change

Answers for Line Master 8**Ionic Compounds—Univalent Metal Ions**

1.

	Name	Formula
a) silver and iodine	silver iodide	$\text{AgI}_{(s)}$
b) magnesium and oxygen	magnesium oxide	$\text{MgO}_{(s)}$
c) magnesium and bromine	magnesium bromide	$\text{MgBr}_{2(s)}$
d) calcium and nitrogen	calcium nitride	$\text{Ca}_3\text{N}_{2(s)}$
e) zinc and selenium	zinc selenide	$\text{ZnSe}_{(s)}$
f) sodium and sulfur	sodium sulfide	$\text{Na}_2\text{S}_{(s)}$
g) barium and phosphorus	barium phosphide	$\text{Ba}_3\text{P}_{2(s)}$
h) aluminium and fluorine	aluminium fluoride	$\text{AlF}_{3(s)}$
i) potassium and chlorine	potassium chloride	$\text{KCl}_{(s)}$
j) silver and oxygen	silver oxide	$\text{Ag}_2\text{O}_{(s)}$

2.

a) MgCl_2	magnesium chloride
b) Ag_3N	silver nitride
c) CsF	cesium fluoride
d) CdO	cadmium oxide
e) MgBr_2	magnesium bromide
f) Al_2O_3	aluminium oxide
g) NaI	sodium iodide
h) K_2S	potassium sulfide
i) BaS	barium sulfide
j) Li_3P	lithium phosphide